

# VISWABHARATHI WISEWOODS

## PRACTICE SHEET

GRADE : VI – VII

SUBJECT : CHEMISTRY

### I. Choose the correct answer :

- Name an element which is common to all acids? ( )  
A) Sulphur      B) Chlorine      C) Nitrogen      D) Hydrogen
- What will be the reaction when acids dissolve in water? ( )  
A) Exothermic      B) Endothermic      C) Substitution      D) Addition
- According to Arrhenius, base is a substance producing \_\_\_\_\_ ion in aqueous solution. ( )  
A)  $H^+$       B)  $OH^-$       C)  $H_3O^+$       D)  $K^+$
- Who gave acid-base theory based on the process of ionization? ( )  
A) Robert Boyle      B) Arrhenius      C) Bronsted-Lowry      D) Lewis
- According to Bronsted-Lowry, an acid is a substance which ( )  
A) accepts proton from another substance      B) donates proton to another substance  
C) does not ionize      D) does not produce  $OH^-$  ion
- According to Bronsted-Lowry, when HCl gas is dissolved in water, water behaves as ( )  
A) an acid      B) a base      C) a neutral substance      D) a salt
- The reaction between an acid and base is called ( )  
A) acid hydrolysis      B) Hydrolysis      C) Alkalisiation      D) Neutralisation
- According to the acid-base theory by Bronsted-Lowry, water is ( )  
A) an acid      B) a base      C) amphoteric      D) neutral
- Which of the following is a strong acid? ( )  
A) Acetic acid      B) Citric acid      C) Nitric acid      D) Oxalic acid
- Which of the following solution is most basic? ( )  
A)  $p^H = 8.2$       B)  $p^H = 9.3$       C)  $p^H = 11.5$       D)  $p^H = 10.6$

11. 20 ml of 10 N HCl and 10 ml of 5 N HCl are mixed and made up to 1 lit with distilled water. The normality of the resulting solution is ( )  
A) 0.25 N                      B) 0.3 N                      C) 0.20 N                      D) 0.1 N
12. What volume of 0.8 M solution contains 0.4 mole of solute? ( )  
A) 100 ml                      B) 125 ml                      C) 500 ml                      D) 62.5 ml
13. Which of the following method of expression concentration is Independent of temperature and have no units? ( )  
A) Molarity                      B) Mole fraction                      C) Molality                      D) Normality
14. A solution whose concentration is exactly known is called ( )  
A) Centimolar solution                      B) Saturated solution  
C) Standard solution                      D) Any of the above
15. The weight of NaOH (in gr) present in 100 ml of 0.5 M NaOH solution is ( )  
A) 1                      B) 3                      C) 2                      D) 4
16. Molarity of pure water (density = 1 gr/ml) is ( )  
A) 40 M                      B) 4 M                      C) 55.6 M                      D) 25 M
17. Molecular weight of  $H_2SO_4$  is ( )  
A) 98                      B) 89                      C) 45                      D) 106
18. Which of the following is correct expression for molarity. ( )  
A)  $M = \frac{n}{V(ml)} \times 1000$                       B)  $M = \frac{V}{n} \times 1000$   
C)  $MV(ml) = n \times 1000$                       D) Both 'A' and 'C'
19. Number of moles present in 23 gr of 'Na' ( )  
A) 0.1                      B) 1                      C) 0.5                      D) 0.01
20. Units for molarity is ( )  
A) mole  $\times$  lit                      B) mole  $\times$  lit<sup>-1</sup>                      C) mole                      D) None of these

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